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web in this experiment you will learn how a hand warmer  
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themselves to design the best all around hand warmer  
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web suggested activity use this resource could be used as a  
whole class investigation with older children observing  
differences and similarities with hand warmers leading to  
research on how they work children can investigate why

materials have been chosen for the hand warmers linking to  
previous knowledge on materials and properties

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web  $m \times c \times \Delta t$  equation 1 where m is the total mass of the  
solution solute plus solvent c is the specific heat of the  
solution and  $\Delta t$  is the observed temperature change the  
specific heat of the solution is generally assumed to be the  
same as that of

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web the heat of reaction  $\Delta H_{\text{soln}}$  is written after the products  
in units of kJ mol rxn for example sodium hydroxide dissolves  
exothermically releasing 44.2 kilojoules per mole dissolved  
this process is represented as  $\text{NaOH (s)} \rightarrow \text{NaOH (aq)}$   $\Delta H_{\text{soln}} = -44.2$   
kJ mol rxn write equations to similarly represent the  
dissolving process for

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web hand warmers are usually placed inside the gloves to  
keep your hands warm the heat keeping your hands warm is  
formed as a result of an exothermic reaction which occurs in  
a small packet hand warmers have 5 main ingredients in  
them salt carbon polypropylene vermiculite water and iron  
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considered dangerous do not touch taste or smell any  
chemicals unless specifically instructed to do so 2 never  
remove chemicals or other materials from the laboratory

area 3 do not eat food drink beverages or chew gum in the laboratory

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web background the hand warmers are filled with sodium acetate click the metal disc and you will notice the liquid changes state and becomes a solid in the process heat is released the sodium acetate is a supersaturated solution *hand warmer chemistry easy chemical hot packs science* - Apr 30 2023

web oct 9 2021 sodium acetate coin or other small piece of metal plastic bags seal the coin inside a small bag place this bag inside a larger bag containing supersaturated sodium acetate get this by dissolving as much sodium acetate as possible in hot water this chemical hand warmer is reusable

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web the goal of my experiment and research is to evaluate the various number of compounds to be used as hand warmers and locate the ones that can raise the temperature to 20 degrees celsius in 100 g water and costs as little as possible

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web in this lab you are challenged to use chemistry to determine which ionic compound makes the best hand warmer in this experiment you will collect data that will allow you to calculate the heat of solution  $\Delta H_{\text{soln}}$  or the amount of energy released or absorbed in kJ/mol of the substance being dissolved

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web mar 4 2014 contains water the hand warmer is struck in a manner that ruptures the inner pouch releasing the ionic salt into the water of the outer pouch the salt dissolves and the water warms this type of hand warmer tends to produce a more vigorous heat than the dry powder type of hand warmer but does not produce heat for quite as long

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the chemistry of crystal formation and heat generation by pressing a button in a pouch which contains a supercooled solution you start a rapid exothermic heat producing crystallization in this science fair project you will determine how the starting temperature affects hand warmer chemistry

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web lab investigation in today s experiment you will design and execute an experimental procedure to determine which of three ionic compounds is most suitable for use in a hand warmer the three compounds you will use is determined by the group number you are assigned the ideal hand warmer